

# Opportunities and Challenges of Industrialization of Hydrogen Production from Natural Seawater

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## Abstract

*Humans have always been looking for sustainable renewable energy, whether from coal, natural gas, hydropower, nuclear energy, solar energy, wind energy, solar thermal energy, hydrogen energy, etc., each energy has its advantages and disadvantages! 2019 Kuang et al. published "Solar-driven, highly sustained splitting of seawater into hydrogen and oxygen fuels" this paper has given the author a great level of research interest in this topic! The researcher tried to explore the challenges and opportunities that this method actually applied to mass production by thinking about various angles.*

Keywords: industrialization, hydrogen, natural seawater, renewable energy

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## I. INTRODUCTION

In recent years, the world's discussion on the issue of hydrogen energy has become more and more heated. The focus of attention is mainly on hydrogen vehicles. Imagine a car without gasoline but hydrogen. The exhaust gas is water, which can also be used for power generation and drinking. This is truly renewable energy! At present, hydrogen is expensive, and every expert and scholar is looking for a suitable method to produce hydrogen at the lowest cost. The traditional process of producing hydrogen from seawater requires purification and then filtering into pure water, which is then decomposed into hydrogen and oxygen. Now, we may be able to separate directly from seawater, which can reduce more processes and costs. This article aims to conduct a preliminary assessment of the direct separation of hydrogen from natural seawater and its current and potential applications in seawater treatment and desalination, with a focus on a rigorous assessment of progress, knowledge gaps, and future research directions from a management perspective.

## II. LITERATURE REVIEW

"Water Will Be the Coal of the Future." (Verne, 1874). Water electrolysis is the most promising method for efficient production of high purity hydrogen (and oxygen), while the required power input for the electrolysis process can be provided by renewable sources (e.g., solar or wind). There are many typical characteristics of the main electrolysis technologies, such as Alkaline electrolysis, Proton Exchange electrolysis & Oxygen ion electrolysis (Sapountzi, F. M. et al. 2017) as per table 1.

Table 1. The typical characteristics of the main electrolysis technologies (Sapountzi, F. M. et al. 2017).

	Low Temperature Electrolysis			High Temperature Electrolysis		
	Alkaline (OH <sup>-</sup> ) electrolysis	Proton Exchange (H <sup>+</sup> ) electrolysis		Oxygen ion(O <sup>2-</sup> ) electrolysis		
	Liquid	Polymer Electrolyte Membrane		Solid Oxide Electrolysis (SOE)		
Operation principles	Conventional	Solid alkaline	H <sup>+</sup> - PEM	H <sup>+</sup> - SOE	O <sup>2-</sup> - SOE	Co-electrolysis
Charge carrier	OH <sup>-</sup>	OH <sup>-</sup>	H <sup>+</sup>	H <sup>+</sup>	O <sup>2-</sup>	O <sup>2-</sup>
Temperature	20-80°C	20-200°C	20-200°C	500-1000°C	500-1000°C	750-900°C
Electrolyte	liquid	solid (polymeric)		solid (ceramic)		
Anodic Reaction (OER)	4OH <sup>-</sup> → 2H <sub>2</sub> O + O <sub>2</sub> + 4e <sup>-</sup>	4OH <sup>-</sup> → 2H <sub>2</sub> O + O <sub>2</sub> + 4e <sup>-</sup>	2H <sub>2</sub> O → 4H <sup>+</sup> + O <sub>2</sub> + 4e <sup>-</sup>	2H <sub>2</sub> O → 4H <sup>+</sup> + 4e <sup>-</sup> + O <sub>2</sub>	O <sup>2-</sup> → 1/2O <sub>2</sub> + 2e <sup>-</sup>	O <sup>2-</sup> → 1/2O <sub>2</sub> + 2e <sup>-</sup>
Anodes	Ni > Co > Fe (oxides) Perovskites: Ba <sub>0.5</sub> Sr <sub>0.5</sub> Co <sub>0.8</sub> Fe <sub>0.2</sub> O <sub>3-δ</sub> , LaCoO <sub>3</sub>	Ni-based	IrO <sub>2</sub> , RuO <sub>2</sub> , Ir <sub>x</sub> Ru <sub>1-x</sub> O <sub>2</sub> Supports: TiO <sub>2</sub> , ITO, TiC	Perovskites with protonic-electronic conductivity	La <sub>0.8</sub> Sr <sub>0.2</sub> MnO <sub>3</sub> + Y-Stabilized ZrO <sub>2</sub> (LSM-YSZ)	La <sub>0.8</sub> Sr <sub>0.2</sub> MnO <sub>3</sub> + Y-Stabilized ZrO <sub>2</sub> (LSM-YSZ)
Cathodic Reaction (HER)	2H <sub>2</sub> O + 4e <sup>-</sup> → 4OH <sup>-</sup> + 2H <sub>2</sub>	2H <sub>2</sub> O + 4e <sup>-</sup> → 4OH <sup>-</sup> + 2H <sub>2</sub>	4H <sup>+</sup> + 4e <sup>-</sup> → 2H <sub>2</sub>	4H <sup>+</sup> + 4e <sup>-</sup> → 2H <sub>2</sub>	H <sub>2</sub> O + 2e <sup>-</sup> → H <sub>2</sub> + O <sup>2-</sup>	H <sub>2</sub> O + 2e <sup>-</sup> → H <sub>2</sub> + O <sup>2-</sup> CO <sub>2</sub> + 2e <sup>-</sup> → CO + O <sup>2-</sup>
Cathodes	Ni alloys	Ni, Ni-Fe, NiFe <sub>2</sub> O <sub>4</sub>	Pt/C MoS <sub>2</sub>	Ni-cermets	Ni-YSZ Subst. LaCrO <sub>3</sub>	Ni-YSZ perovskites
Efficiency	59-70%		65-82%	up to 100%	up to 100%	-
Applicability	commercial	laboratory scale	near-term commercialization	laboratory scale	demonstration	laboratory scale
Advantages	low capital cost, relatively stable, mature technology	combination of alkaline and H <sup>+</sup> -PEM electrolysis	compact design, fast response/start-up, high-purity H <sub>2</sub>	enhanced kinetics, thermodynamics: lower energy demands, low capital cost		+ direct production of syngas
Disadvantages	corrosive electrolyte, gas permeation, slow dynamics	low OH <sup>-</sup> conductivity in polymeric membranes	high cost polymeric membranes; acidic: noble metals	mechanically unstable electrodes (cracking), safety issues: improper sealing		
Challenges	Improve durability/reliability; and Oxygen Evolution	Improve electrolyte	Reduce noble-metal utilization	microstructural changes in the electrodes: delamination, blocking of TPBs, passivation		C deposition, microstructural change electrodes

Bureau of Energy, Ministry of Economic Affairs, Republic of China (Taiwan) listed the importance of hydrogen energy in the 2007 white paper on energy technology research and development, and also proposed various possible ways and methods of hydrogen energy production. It also mentioned the use of renewable energy to electrolyze water to produce hydrogen, and this method is suitable for seawater and freshwater. Various possible ways and methods of hydrogen energy production are shown in Figure 1.

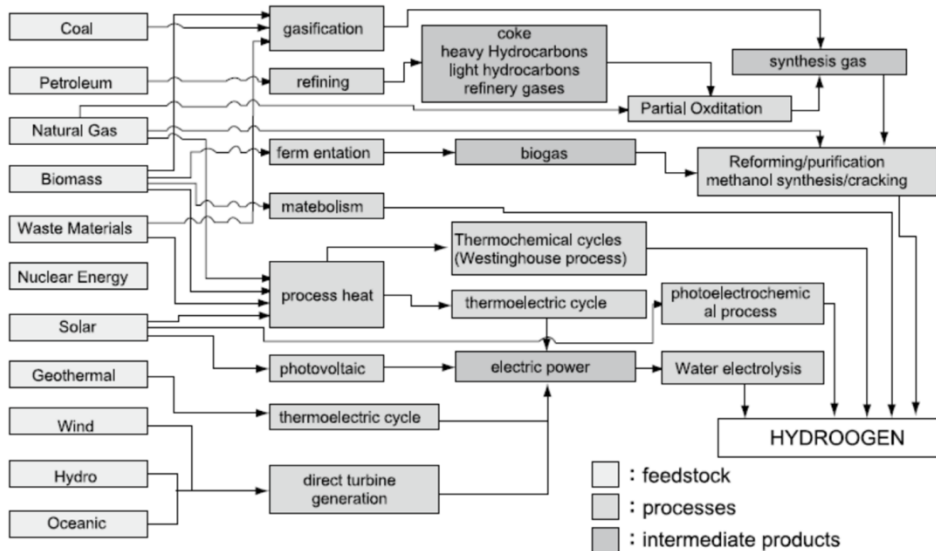


Figure 1. Various possible ways and methods of hydrogen energy production (BOE, 2007)

For example, the Offshore wind power produces hydrogen as per Figure 2.

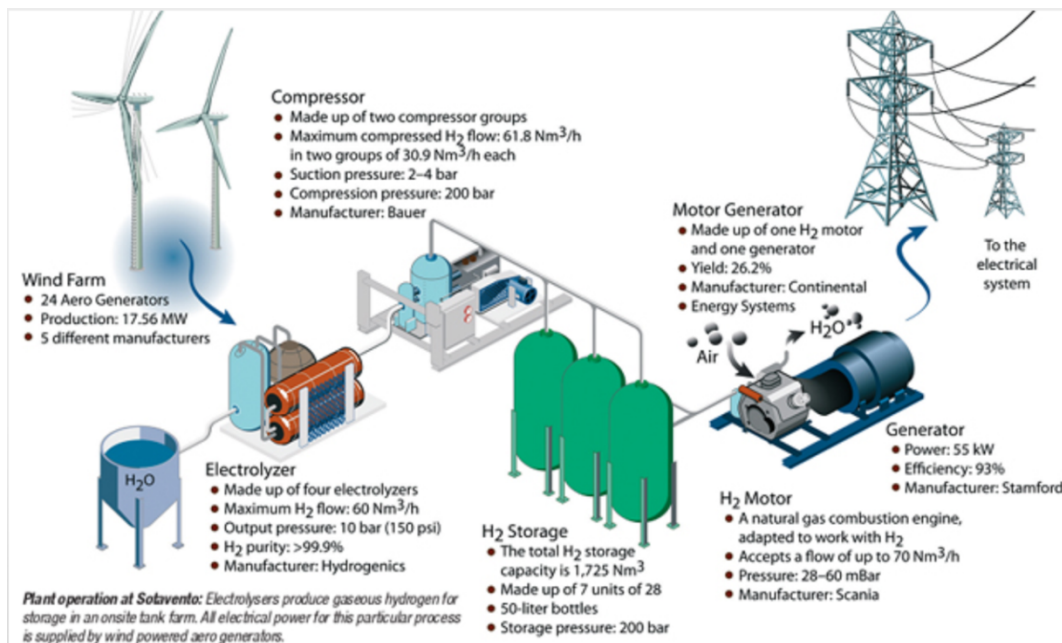


Figure 2. Offshore wind power produces hydrogen - Sotavento hydrogen-wind power plant. (Crump, 2010).

### II.1 Solar electrolysis of seawater to produce hydrogen

When the water electrolysis hydrogen production, the anode should completely avoid poisonous chlorine precipitation. Using traditional anode electrolytic seawater electrolysis of water use, with chlorine precipitation is given priority to, to avoid chlorine corrosion of equipment and the environmental pollution, the development of a chlorine oxygen inhibition anode materials is the key

of electrolytic water hydrogen plot development technology (Zhang et al., 2009). In 2014, Xiamen University applied for a utility model patent and disclosed a device for the electrolysis of seawater by solar energy to produce hydrogen as per figure 3, which includes an electrolytic cell, an ion exchange membrane, a cathode gas collection device, a sheep pole gas collection device, and a nuclear electrolysis shop level. It also includes a seawater filtration device and solar energy. Photovoltaic cells, seawater used for electrolytic hydrogen production in electrolytic cells, seawater filtering devices are passed through the electrolytic cells, the seawater has been filtered and introduced into the electrolytic cells; a solar photovoltaic cell store is connected to a transformer, and the transformer is connected to the anode and cathode electrolytic power through the wire. The seawater is used as the raw material for hydrogen electrolysis, with abundant seawater resources, wide sources, low production costs, clean and pollution-free, and can also bring additional products during the electrolysis process; the solar energy is converted into electrical energy through solar photovoltaic cells to the electrolytic power. Compared with the traditional electric power supply, it has a lower cost, sufficient power supply, pure natural, and pollution-free advantages. (CN 203976930)

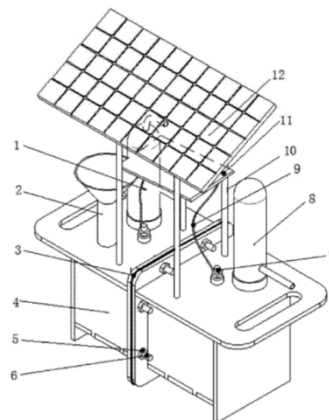


Figure 3. Solar electrolysis of seawater to produce hydrogen (CN 203976930).

## II.2 Photocatalytic hydrogen production

The Photocatalyst effect is shown in Figure 4., also known as the "Honda-Fujishima effect," which was discovered by two scholars, Japan's Bunta Kenichi and Fujishima Akira. In 1967, Professor Kenichi Bunta and his graduate student Akira Fujishima discovered that using titanium dioxide and platinum as electrodes, placing them in water and irradiating them with light, even if they were not powered, they could decompose water into oxygen and hydrogen.

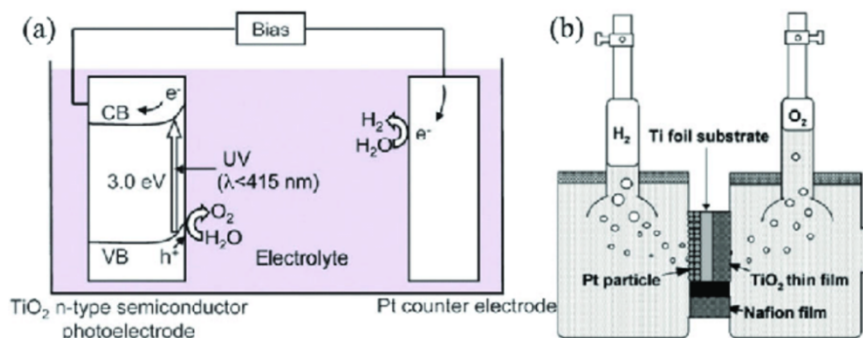


Figure 4. The Photocatalyst effect (Fujishima et al., 1972).

### II.3 Hydrogen energy from the ocean

nanomaterials make photocatalysts produce hydrogen more efficiently, as per figure 5. Researchers have found a way to make photocatalyst composites to solve this efficiency problem. They found that the surface of titanium dioxide was chemically "etched" into the pores of the nanostructure. A nanofilm with a thickness of only one atom is laid on the surface: molybdenum disulfide by adjusting the vacancy density of sulfur atoms in the surface film. Enables photocatalysts to accept larger photon bandwidths: from ultraviolet-visible to near-infrared light, The efficiency of photocatalytic hydrolysis is nearly doubled. (Guo et al. 2018)

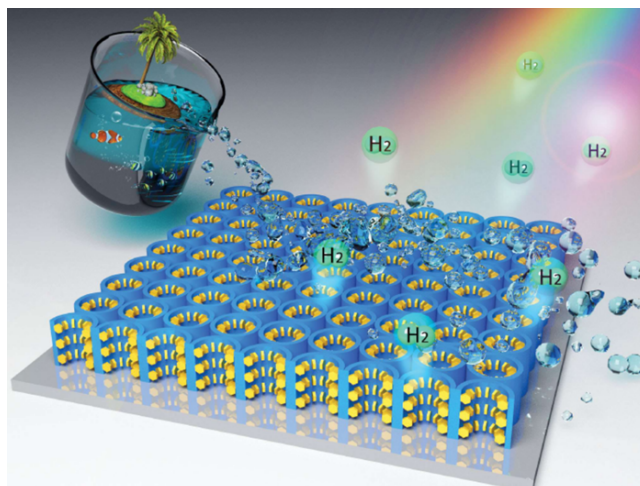


Figure 5. nanomaterials make photocatalysts produce hydrogen more efficiently (Guo et al. 2018).

### II.4 Photocatalytic hydrogen production

In 2018, Li Can, a member of the Chinese Academy of Sciences' Solar Energy Research Department of the Chinese Academy of Sciences and researcher Zhang Fuxiang, made new progress in the study of the wide-spectrum catalyst Z mechanism for the complete decomposition of water and hydrogen. The results were published in the Joule Journal of Cell. (Qi et al., 2018) are shown in Figure 6.

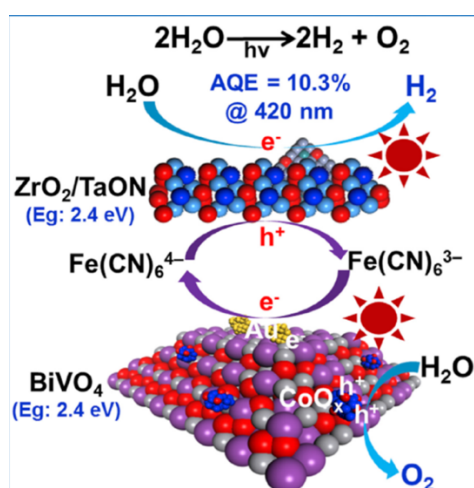


Figure 6. wide-spectrum catalyst Z mechanism for the complete decomposition of water and hydrogen (Qi et al, 2018).

### II.5 Micro-nano structured electrode electrolysis of seawater to produce hydrogen

Ni—Fe—C alloys have also been used for seawater electrolysis, and it was found that their activity depends on the grain size and the composition of the alloy (Sapountzi et al. 2017). Electrolysis of water to produce hydrogen fuel is an attractive renewable energy storage technology. However, grid-scale electrolysis of freshwater can put heavy pressure on important water resources. The development of cheap electrocatalysts and electrodes to maintain cracking without chloride corrosion could solve water scarcity problems. A multilayer anode composed of a nickel-iron hydroxide (NiFe) electrocatalyst layer uniformly distributed on a nickel sulfide (NiS<sub>x</sub>) layer formed on a porous nickel foam (NiFe / NiS<sub>x</sub> -Ni), which has excellent catalytic activity and Corrosion-resistant alkaline seawater electrolysis operates at 1,000 hrs of industrially required current density (0.4 to 1 A / cm<sup>2</sup>) are shown in Figure 7. A continuous, highly oxygen evolution reaction-active NiFe electrocatalyst layer drawing anodic currents toward water oxidation and an in situ-generated polyatomic sulfate and carbonate-rich passivating layers formed in the anode are responsible for chloride repelling and superior corrosion resistance of the salty-water-splitting anode. (Kuang et al. 2019)

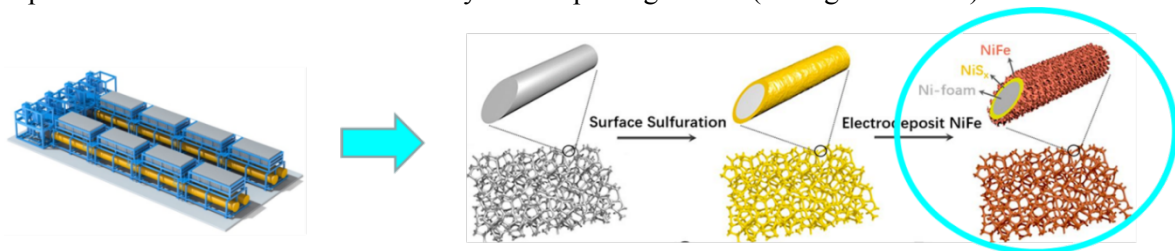


Figure 7. Micro-nano structured electrode (Kuang et al. 2019).

### II.6 Golden Gate Zero Emission Marine (GGZEM)

The vessel has been dubbed the "water-go-round," a vivid illustration of how fuel cell technology works in cycles. Specifically, hydrogen can be generated from water, and then water can be generated after fuel cell use, truly achieving zero emissions and pollution-free. The all-aluminum passenger ship, designed by Incat Crowther, is 70 feet long and can seat 84 passengers at a top speed of 22 knots. (Pratt, J. W., & Burgard, J. 2016) as per Figure 8.

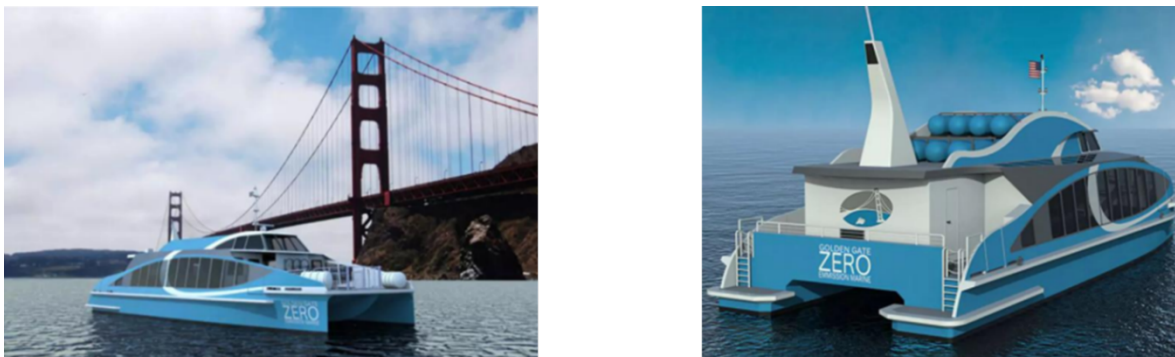


Figure 8. Golden Gate Zero Emission Marine (Pratt, J. W., & Burgard, J. 2016).

### III. RESEARCH METHODOLOGY

According to the "Solar-driven, highly sustained splitting of seawater into hydrogen and oxygen fuels" thesis, the new process has been studied on an experimental scale, and the hydrogen can be directly produced by electrolysis of seawater through micro-nano structured electrodes without considering the desalination process and has a stability of thousands of hours under the current density of industrial electrolysis. This study intends to consolidate relevant literature and existing processes on seawater hydrogen production and find comparative studies based on the opportunity and challenge framework.

**Document Analysis:** Refers to a method that comprehensively and accurately grasps the problem to be researched by collecting relevant market information, survey reports, industry trends, and other literature data according to a certain research purpose or topic. The content of the collection should be as rich and extensive as possible, and then the data collected around should be summarized after analysis, and then the source, cause, background, impact, and significance of the event should be analyzed. Documents can be reports from government departments, research from the business sector, documentary databases, corporate organization materials, books, papers, and periodicals in the library, newspapers, and more. There are four analysis steps, namely Reading and Organizing, Description, Classifying, and Interpretation (Titscher et al. 2000)

### IV. FINDING AND DISCUSSION

Identified opportunities and challenges are as per table 2.

Table 2. Opportunities and challenges.

Item	Years	Method	Details	Opportunity	Challenge	Remark
1	2014	<b>Solar hydrogen production from seawater electrolysis</b>	The seawater will be <b>filtered and discharged</b> into the electrolytic cell. The solar photovoltaic cells are connected to a transformer, which is connected to the anode and cathode via wires. In this method, seawater is used as a raw material for hydrogen production	Solar photovoltaic cells convert <b>solar energy into electrical energy</b> to power electrolytic stages, which has a lower cost than traditional electrical power.	No sunlight at night and rainy days cannot convert light energy into electricity	Xiamen University (CN 203976930)
2	2016	<b>Seawater electrolyzer system</b>	Large-scale systems: (1) seawater pressurization system, (2) seawater thorium system, (3) seawater electrolysis system, (4) <b>hydrogen release system</b> , (5) sodium hypochlorite storage and injection system, and (6) pickling system.	The existing seawater electrolysis equipment has a hydrogen release system, which can be transformed into a hydrogen collection system by reference.	The anode produces a lot of chlorine instead of oxygen (highly toxic and highly corrosive)	Mainly depends on technology type and load factor (TW201738460A)
3	2018	<b>Photocatalytic hydrogen production</b>	Nanotechnology has produced high-efficiency titanium dioxide photocatalysts, which can produce hydrogen from photosynthesis in the ultraviolet to near-infrared range covering the entire visible light band. Each gram of catalyst can produce 0.58 moles of hydrogen per hour. The efficiency is the existing More than 2 times of other catalysts, it also has the advantage of resistance to seawater corrosion.	The heterostructure of non-metallic plasma molybdenum disulfide in titanium dioxide for efficient photocatalytic hydrogen production is a promising material for seawater hydrogen production.	Seawater corrosion resistance need more evidence.	(Guo, L. et al. 2018)

4	2019	<b>Solar-driven, highly sustainable seawater breakdown into hydrogen and oxygen fuels</b>	When the pH of seawater is adjusted to alkaline, the oxidation of chloride ions can be suppressed, so that oxygen is more easily generated at the anode	1. The generated oxygen can be effectively used.	1. The salinity and density of seawater in each area are different. How to define alkaline conditions? 2. As seawater pollutants have not been removed, the use of electrodes affected by potential problems of attachment, growth and reproduction of sea creatures in the electrodes.	(Kuang et al. 2019)
			Under alkaline conditions, nickel-iron hydroxide is currently the highest performing and most stable catalyst for the evolution of oxygen.	1. Nickel-iron hydroxide catalyst has advantages in cost.	1. Will the uniformity of catalyst coating of nickel-iron hydroxide not produce oxygen at the anode and may generate sodium hypochlorite?	
			Nickel sulfide is grown on a foamed nickel conductor, and a nickel-iron hydroxide catalyst is grown on top of the nickel sulfide to form a multilayer structure	1. Save pure water (fresh water).	1. Is the electrodeposited multilayer structure stable? Have you tried using physical meteorological deposition or even chemical meteorological deposition to make electrodes?	
					2. Which preparation method is relatively stable and not easily corroded by chloride ions? Is the entire hydrogen production process safe?	

## V. CONCLUSION AND FURTHER RESEARCH

With the continuous advancement of science and technology, we must continue to pay attention to the development of new technologies. At present, there has been a major breakthrough in the direct conversion of seawater to hydrogen and oxygen. However, we still have to measure the input costs and benefits from the perspective of management! At this stage, the use of hydrogen energy is not comprehensive. The production, storage, transportation, and application of hydrogen are still dangerous. In particular, the existing seawater electrolysis mainly involves thermal power plants drawing seawater as cooling water to cool and generate electricity. The residual heat of the sea and the marine attached organisms will enter the cooling system to adhere, grow, and multiply. This will not only reduce the efficiency of the heat exchanger but also cause blockage of the pipeline, obstruct the flow of water and cause corrosion and perforation of the pipeline. In severe cases, it will even force the unit to shut down and affect its operation. Therefore, seawater electrolysis equipment is used to produce sodium hypochlorite to inhibit and prevent the growth of marine organisms in the cooling water system pipeline. However, the application method of seawater electrolysis assigns hydrogen to flammable gas and is prone to explosion. Therefore, hydrogen is separated from seawater through a centrifugal gas-water separator and discharged to the atmosphere, instead of collecting and storing hydrogen as an energy source.

A more reasonable way is to optimize and retrofit the existing seawater electrolysis, collect the currently generated hydrogen reasonably, and gradually experiment with new technologies in small-scale tests and apply them to the existing seawater electrolysis equipment, and install a hydrogen power generation system. The generated electricity is connected to the existing power grid!

Further research is still collecting more efficient, safe, and low-cost methods; it is best to combine the existing seawater electrolysis module to modify or optimize the method, regardless of the improvement of electrode materials, cost reduction, and replacement efficiency. The most important issue is how to effectively reduce the risk of explosion in the production, storage, transportation, and application of hydrogen, and then convert the inexhaustible and inexhaustible seawater into hydrogen. As an energy source, as another effective application of oxygen, we will get rid of the hidden worry of energy depletion.



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